

Hewlett-Packard Professional Certification

Examination 1017 has six to eight following problems. Use all appropriate and proper scientific units. Use scientific notation for answers requiring simplification of the number 1000.

1. Find the standard enthalpy change for the reaction $\text{C}_{10}\text{H}_8 + 12\text{H}_2\text{O}_{(l)} \rightarrow 10\text{H}_2\text{O}_{(g)}$
 of element data: $\text{C}_{10}\text{H}_8 + 8\text{H}_2\text{O}_{(l)} \rightarrow 10\text{H}_2\text{O}_{(g)}$ $\Delta H^\circ = -2000 \text{ kJ}$
 and $\text{C}_{10}\text{H}_8 + 8\text{H}_2\text{O}_{(l)} \rightarrow 10\text{H}_2\text{O}_{(g)} + 1000 \text{ kJ}$

a. $\Delta H^\circ =$ $\text{C}_{10}\text{H}_8 + 12\text{H}_2\text{O}_{(l)} \rightarrow 10\text{H}_2\text{O}_{(g)}$ $\Delta H^\circ =$ $\text{C}_{10}\text{H}_8 + 12\text{H}_2\text{O}_{(l)} \rightarrow 10\text{H}_2\text{O}_{(g)}$

b. $\Delta H^\circ =$ $\text{C}_{10}\text{H}_8 + 12\text{H}_2\text{O}_{(l)} \rightarrow 10\text{H}_2\text{O}_{(g)}$ $\Delta H^\circ =$ $\text{C}_{10}\text{H}_8 + 12\text{H}_2\text{O}_{(l)} \rightarrow 10\text{H}_2\text{O}_{(g)}$

2. The standard enthalpy change for the formation of aluminum oxide and aluminum nitride are -1676 kJ/mol and -1792 kJ/mol respectively. $\Delta H^\circ = -1676 \text{ kJ}$
 $\text{Al}_2\text{O}_3 + 2\text{N}_2\text{O}_{(g)} \rightarrow 2\text{AlN}_{(s)} + 3\text{O}_{2(g)}$ $\Delta H^\circ = -1792 \text{ kJ}$
 Calculate the standard enthalpy of formation: $\text{Al}_2\text{O}_3 + 2\text{N}_2\text{O}_{(g)} \rightarrow 2\text{AlN}_{(s)} + 3\text{O}_{2(g)}$

a. $\Delta H^\circ =$ $\text{Al}_2\text{O}_3 + 2\text{N}_2\text{O}_{(g)} \rightarrow 2\text{AlN}_{(s)} + 3\text{O}_{2(g)}$ $\Delta H^\circ =$ $\text{Al}_2\text{O}_3 + 2\text{N}_2\text{O}_{(g)} \rightarrow 2\text{AlN}_{(s)} + 3\text{O}_{2(g)}$

b. $\Delta H^\circ =$ $\text{Al}_2\text{O}_3 + 2\text{N}_2\text{O}_{(g)} \rightarrow 2\text{AlN}_{(s)} + 3\text{O}_{2(g)}$ $\Delta H^\circ =$ $\text{Al}_2\text{O}_3 + 2\text{N}_2\text{O}_{(g)} \rightarrow 2\text{AlN}_{(s)} + 3\text{O}_{2(g)}$

c. $\Delta H^\circ =$ $\text{Al}_2\text{O}_3 + 2\text{N}_2\text{O}_{(g)} \rightarrow 2\text{AlN}_{(s)} + 3\text{O}_{2(g)}$ $\Delta H^\circ =$ $\text{Al}_2\text{O}_3 + 2\text{N}_2\text{O}_{(g)} \rightarrow 2\text{AlN}_{(s)} + 3\text{O}_{2(g)}$

3. Find the Gibbs free energy for the reaction of carbon monoxide and hydrogen, which was given in question 1016: $\text{CO}_{(g)} + \text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)} + \text{CO}_{(g)}$ $\Delta G^\circ = 71$

Calculate the standard enthalpy change for the reaction from the following standard enthalpies:

$\text{CO}_{(g)} + \text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)} + \text{CO}_{(g)}$ $\Delta H^\circ =$ $\text{CO}_{(g)} + \text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)} + \text{CO}_{(g)}$
 $\text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)}$ $\Delta H^\circ =$ $\text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)}$

a. $\Delta H^\circ =$ $\text{CO}_{(g)} + \text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)} + \text{CO}_{(g)}$ $\Delta H^\circ =$ $\text{CO}_{(g)} + \text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)} + \text{CO}_{(g)}$
 $\text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)}$ $\Delta H^\circ =$ $\text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)}$

b. $\Delta H^\circ =$ $\text{CO}_{(g)} + \text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)} + \text{CO}_{(g)}$ $\Delta H^\circ =$ $\text{CO}_{(g)} + \text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)} + \text{CO}_{(g)}$
 $\text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)}$ $\Delta H^\circ =$ $\text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)}$

c. $\Delta H^\circ =$ $\text{CO}_{(g)} + \text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)} + \text{CO}_{(g)}$ $\Delta H^\circ =$ $\text{CO}_{(g)} + \text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)} + \text{CO}_{(g)}$

Hess Law Problems With Answers

David Baud



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Table of Contents Hess Law Problems With Answers

1. Understanding the eBook Hess Law Problems With Answers
 - The Rise of Digital Reading Hess Law Problems With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Hess Law Problems With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hess Law Problems With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hess Law Problems With Answers
 - Personalized Recommendations
 - Hess Law Problems With Answers User Reviews and Ratings

- Hess Law Problems With Answers and Bestseller Lists
- 5. Accessing Hess Law Problems With Answers Free and Paid eBooks
 - Hess Law Problems With Answers Public Domain eBooks
 - Hess Law Problems With Answers eBook Subscription Services
 - Hess Law Problems With Answers Budget-Friendly Options
- 6. Navigating Hess Law Problems With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Hess Law Problems With Answers Compatibility with Devices
 - Hess Law Problems With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hess Law Problems With Answers
 - Highlighting and Note-Taking Hess Law Problems With Answers
 - Interactive Elements Hess Law Problems With Answers
- 8. Staying Engaged with Hess Law Problems With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hess Law Problems With Answers
- 9. Balancing eBooks and Physical Books Hess Law Problems With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hess Law Problems With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Hess Law Problems With Answers
 - Setting Reading Goals Hess Law Problems With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hess Law Problems With Answers
 - Fact-Checking eBook Content of Hess Law Problems With Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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