

12-2 Study Guide and Intervention

Angles and Angle Measure

Angles in Standard Position An angle is determined by two rays. The degree measure of an angle in standard position is described by the amount and direction of rotation from the **initial side**, which lies along the positive x -axis, to the **terminal side**. A counterclockwise rotation is associated with positive angle measure and a clockwise rotation is associated with negative angle measure. Two or more angles in standard position with the same terminal side are called **coterminal angles**.

Example 1: Draw an angle with measure 290° in standard position.

The negative y -axis represents a positive rotation of 270° . To generate an angle of 290° , rotate the terminal side 20° more in the counterclockwise direction



Example 2: Find an angle with a positive measure and an angle with a negative measure that are coterminal with each angle.

a. 250°

A positive angle is $250^\circ + 360^\circ$ or 610° . Add 360° .
A negative angle is $250^\circ - 360^\circ$ or -110° . Subtract 360° .

b. -140°

A positive angle is $-140^\circ + 360^\circ$ or 220° . Add 360° .
A negative angle is $-140^\circ - 360^\circ$ or -500° . Subtract 360° .

Exercises

Draw an angle with the given measure in standard position.

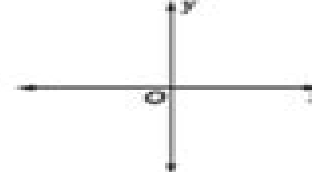
1. 160°



2. 280°



3. 400°



Find an angle with a positive measure and an angle with a negative measure that are coterminal with each angle.

4. 65°

5. -75°

6. 230°

7. 420°

Geometry Study Guide Intervention Work Answers

Judd E. Hollander



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each student boiling point is the point at which a liquid changes physical states and becomes a vapor or gas boiling point is
important to understand because it can be used to help distinguish between different substances

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the boiling point of an organic liquid is the temperature at which its vapour pressure equals the atmospheric pressure over
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point is considered as a criterion of purity of a compound and is useful for identification of

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given that the boiling points of ethanol and methanol 78 c and 64 c respectively are far off in a significant way this seems like
a valid conclusion conclusion thermometer calibrations are necessary in order to collect reliable information as precise
calculations are imperative to gathering correct results

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melting point 1 a push the open end of a capillary tube into the powdered acetamide b move the powder to the closed end of
the capillary tube by tapping it on the table repeat until the the powdered acetamide occupies 1 2 mm of the capillary tube

end c

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boiling point determination lab report conclusion objective to correctly measure boiling issues of various liquids and identify an unknown liquid based on its boiling point background information objects can be distinguished into acid properties and physical properties

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discussion and conclusion the technique recrystallization was used because it is a simple method to purify a solid compound the percentage yield for benzoic acid and acetanilide acid being high indicated the initial sample had a low concentration of impurities with the melting point being in a specific

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determination of boiling points introduction the boiling point of a compound is the temperature at which it changes from a liquid to a gas this is a physical property often used to identify substances or to check the purity of the compound it

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clean up solid and liquid spills immediately true conclusions after conducting the boiling point determination lab it can be concluded that the unknown substance 193 is isopropyl alcohol on account of its observed boiling point of 88 c

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conclusion after completing this experiment the melting point of three different unknown compounds was obtained using the melting point apparatus compound a had a melting point range of 114.6-115.5 c thus revealing it to be acetanilide compound b had a melting point range of 122.6-124.5 c thus

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melting point can be used as the identity and purity of an organic solid however it is not the same as the boiling point the boiling point is a characteristic unique of pure liquids the boiling point is the temperature at which the vapor pressure of a liquid is equal to the atmospheric pressure

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50 °C so in order to convert an experimental measurement taken in Calgary which is at higher altitude than sea level and so is at a lower pressure to that reported for sea level higher pressure one needs to add a correction factor since the boiling point at sea level is higher than that at higher altitudes

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introduction in chemistry one may encounter an unknown substance that needs to be determined by physical properties such as boiling point melting point density solubility etc in this experiment I aim to figure out the boiling point of a mystery substance in order to identify it chemically this experiment requires the handling of a Bunsen burner boiling

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there are a variety of methods by which a sample's boiling point can be determined including distillation reflux and by using a Thiele tube the most straightforward method uses a Thiele tube and has the advantage of using less than 0.5 mL of material

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