

**12.2****FACTORS AFFECTING GAS PRESSURE****SECTION REVIEW****Objectives**

- Explain how the amount of gas and the volume of the container affect gas pressure.
- Infer the effect of temperature changes on the pressure exerted by a contained gas.

**Part A Completion**

Use this completion exercise to check your understanding of the concepts and terms a chapter has introduced. In each sentence, insert the word or phrase which is correct, absent, or incorrect.

- As long as the temperature of an enclosed gas does not change, \_\_\_\_\_ 1. \_\_\_\_\_, doubling the number of gas particles \_\_\_\_\_ 2. \_\_\_\_\_ the pressure. When gas particles are removed from a container the pressure decreases. \_\_\_\_\_ 3. \_\_\_\_\_
- Reducing half the gas particles decreases the pressure by \_\_\_\_\_ 4. \_\_\_\_\_.
- When the number of gas particles remains constant, the \_\_\_\_\_ 5. \_\_\_\_\_ pressure can be increased by \_\_\_\_\_ 6. \_\_\_\_\_ the volume. Reducing the volume of a contained gas by half \_\_\_\_\_ 7. \_\_\_\_\_ the pressure.
- When the number of gas particles and the volume of a gas remain constant, the pressure of a gas can be increased by increasing the \_\_\_\_\_ 8. \_\_\_\_\_. If the Kelvin temperature doubles, the pressure of the contained gas \_\_\_\_\_ 9. \_\_\_\_\_. If the pressure decreases by half when the Kelvin temperature is halved.

**Part B True-False**

Classify each of these statements as true always true, AT; sometimes true, ST; or never true, NT.

- \_\_\_\_\_ 1. Halving the number of particles in a given volume of gas decreases the pressure by one-half if the temperature is kept constant.
- \_\_\_\_\_ 2. Reducing the volume of a container of gas by one-half has the same effect on pressure as doubling the quantity of the gas in the container by one-half, if temperature is kept constant.
- \_\_\_\_\_ 3. When a sealed container of gas under pressure is opened, gas particles move from the container to the vacuum of air.

# Gases And Pressure Section Review Answers

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