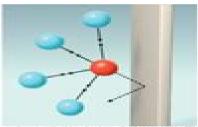
## Chapter 11: Gases

## Focus of this section:

## Deviation from ideality

## Deviation from Ideal Behavior:

- recall: real gases deviate from ideal gas behavior at high pressure and low temperature conditions
  - 1. real molecules have volume
    - -- cannot compress gas to zero volume
  - 2. real molecules interact (attracted to each other)
    - causes lower P (vs. ideal gas), especially at lower temperatures



- = Since 2 molecules cannot occupy the same space, V<sub>accessible</sub> < V<sub>container</sub> and a corrective term should be <u>subtracted</u> from V in ideal gas equation
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experimentally measured pressure 
$$\left(P + \frac{an^2}{V^2}\right)(V - nb) = nRT$$

$$P_{\text{ideal}} = \left(P_{\text{obs}} + \frac{n^2 a}{V^2}\right)$$
 correction  $\propto \left(\frac{n}{V}\right)^2$ 
 $V_{\text{ideal}} = \left(V_{\text{container}} - nb\right)$  correction  $\propto n$ 

PV = nRT

= Note, "a" and "b" are different for each gas

# **Gases Section Quizzes**

Christopher R. Hren, John T. Moore, Peter J. Mikulecky

#### **Gases Section Quizzes:**

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