

7. If you have 17 liters of gas at a temperature of 67 °C and a pressure of 88.89 atm, what will be the pressure of the gas if you raise the temperature to 94 °C and decrease the volume to 12 liters?

$$P_2 = \frac{P_1 V_1 T_2}{V_2 T_1} = \frac{88.89 \text{ atm} (17 \text{ L}) 367 \text{ K}}{(12 \text{ L}) 340 \text{ K}}$$

$$P_2 = \quad \text{atm}$$

8. I have an unknown volume of gas at a pressure of 0.5 atm and a temperature of 325 K. If I raise the pressure to 1.2 atm, decrease the temperature to 320 K, and measure the final volume to be 48 liters, what was the initial volume of the gas?

$$V_1 = \frac{P_2 V_2 T_1}{T_2 P_1} = \frac{1.2 \text{ atm} (48 \text{ L}) 325 \text{ K}}{320 \text{ K} (0.5 \text{ atm})}$$

$$V_1 = \quad \text{L}$$

9. If you have 21 liters of gas held at a pressure of 78 atm and a temperature of 900 K, what will be the volume of gas if you decrease the pressure to 45 atm and decrease the temperature to 750 K?

$$V_2 = \frac{78 \text{ atm} (21 \text{ L}) 750 \text{ K}}{900 \text{ K} (45 \text{ atm})}$$

$$V_2 = \quad \text{L}$$

10. If I have 2.9 L of gas at a pressure of 5 atm and a temperature of 50 °C, what will the temperature of the gas be if I decrease the volume of the gas to 2.4 L and decrease the pressure to 3 atm?

$$1/T_2 = \frac{P_1 V_1}{T_1 P_2 V_2} = \frac{(2.9 \text{ L}) 5 \text{ atm}}{323 \text{ K} (3 \text{ atm}) 2.4 \text{ L}}$$

$$T_2 = 160 \text{ K}$$

11. You have an unknown volume of gas held at a temperature of 115 K in a container with a pressure of 60 atm. If by increasing the temperature to 225 K and decreasing the pressure to 30 atm causes the volume of the gas to be 29 L, how many liters of gas did you start with?

$$V_1 = \frac{P_2 V_2 T_1}{T_2 P_1} = \frac{115 \text{ K} (29 \text{ L}) 30 \text{ atm}}{(225 \text{ K}) 60 \text{ atm}}$$

$$V_1 = \quad \text{L}$$

Gas Laws Practice Calculations Key

Michael E. Browne

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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