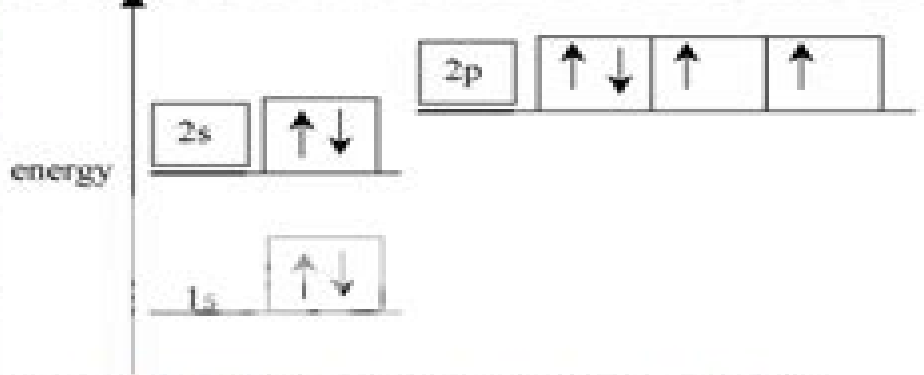


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Question	Expected answers	Marks
1 (a)	  ✓ 1, 2 or 3 p orbitals are OK	[2]
(b)	d orbital 2 ✓ p sub-shell 6 ✓ 3rd shell 18 ✓	[3]
(c) (i) (ii)	 2s and 2p labels ✓ Ignore any superscripted numbers. 8 electrons in correct levels with arrows correctly shown ✓	[2]
(d) (i) (ii)	(i) Energy change when each atom in 1 mole ✓ of gaseous atoms ✓ loses an electron ✓ (to form 1 mole of gaseous 1+ ions). (ii) $O^{2+}(g) \longrightarrow O^{3+}(g) + e^{-}$ ✓✓ 1 mark for correct species; 1 mark for state symbols No charge required on electron. Ignore (g) on e	[3] [2]

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Table of Contents F321 Past Papers

1. Understanding the eBook F321 Past Papers
 - The Rise of Digital Reading F321 Past Papers
 - Advantages of eBooks Over Traditional Books
2. Identifying F321 Past Papers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an F321 Past Papers
 - User-Friendly Interface
4. Exploring eBook Recommendations from F321 Past Papers
 - Personalized Recommendations
 - F321 Past Papers User Reviews and Ratings
 - F321 Past Papers and Bestseller Lists
5. Accessing F321 Past Papers Free and Paid eBooks
 - F321 Past Papers Public Domain eBooks
 - F321 Past Papers eBook Subscription Services

-
- F321 Past Papers Budget-Friendly Options
6. Navigating F321 Past Papers eBook Formats
 - ePub, PDF, MOBI, and More
 - F321 Past Papers Compatibility with Devices
 - F321 Past Papers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of F321 Past Papers
 - Highlighting and Note-Taking F321 Past Papers
 - Interactive Elements F321 Past Papers
 8. Staying Engaged with F321 Past Papers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers F321 Past Papers
 9. Balancing eBooks and Physical Books F321 Past Papers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection F321 Past Papers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine F321 Past Papers
 - Setting Reading Goals F321 Past Papers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of F321 Past Papers
 - Fact-Checking eBook Content of F321 Past Papers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends

-
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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