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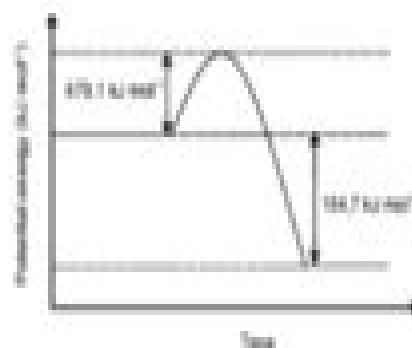
# CHEMICAL CHANGE

# 11

## ENDOTHERMIC & EXOTHERMIC REACTIONS

QUESTION 1 (start on a new page)

The diagram shows the potential energy changes during the following chemical reaction:



- 1.1 Define activation energy. (2)
- 1.2 Is the reaction  $2\text{H}_2\text{g} + \text{O}_2\text{g} \rightarrow 2\text{H}_2\text{Og}$  endo/exothermic? (1)  
Give a reason for the answer. (2)
- 1.3 What is the total bond energy ( $\Sigma E_{\text{b}}$ ) of the reactants? (1)  
Give a reason for the answer. (2)
- 1.4 Determine the energy released by the bond formation of the  $\text{H}_2\text{O}$  molecule. (2)
- 1.5 What effect will the addition of a catalyst have on the value 166.7 kJ mol<sup>-1</sup>? (1)  
Write down only INCREASE, DECREASE or NO EFFECT. (1)  
Give a reason for the answer. (2)

# Endothermic Investigation Grade 11 2015

**Forman A. Williams**



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