

Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.0 g of carbon and 6.67 g of hydrogen. Find the empirical formula.

$$\text{C}_{40.0} \text{H}_{6.67} \times \frac{1 \text{ mole}}{12.01 \text{ g C}} = 3.33 \text{ moles C}$$

$$\text{H}_{6.67} \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 6.62 \text{ moles H}$$

C₃H₅

2. A compound is found to contain 50.0 g of sulfur, 40.0 g of oxygen, and 10.0 g of hydrogen. Find the empirical formula.

$$\text{S}_{50.0} \text{O}_{40.0} \text{H}_{10.0} \times \frac{1 \text{ mole}}{32.07 \text{ g S}} = 1.56 \text{ moles S}$$

$$\text{O}_{40.0} \times \frac{1 \text{ mole}}{16.00 \text{ g O}} = 2.50 \text{ moles O}$$



$$\text{H}_{10.0} \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 9.92 \text{ moles H} \rightarrow \text{C}_2 \text{H}_2 \text{O}_5 \text{S}$$

3. A 10.0 g sample of a compound, a possible alkane, is found to contain 8.56 g of carbon, 1.44 g of hydrogen, and 0.00 g of oxygen. Find the empirical formula.

$$\text{C}_{8.56} \text{H}_{1.44} \times \frac{1 \text{ mole}}{12.01 \text{ g C}} = 0.713 \text{ moles C}$$



$$\text{H}_{1.44} \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 1.43 \text{ moles H} \rightarrow \text{C}_5 \text{H}_8$$

$$\text{C}_{1.74} \text{H}_1 \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 1.72 \text{ moles C} \rightarrow \text{C}_5 \text{H}_8$$

4. A 20.00 g sample of an unknown compound contains 12.00 g carbon, 8.00 g hydrogen, and 0.00 g oxygen. Find the empirical formula.

$$\text{C}_{12.00} \text{H}_{8.00} \times \frac{1 \text{ mole}}{12.01 \text{ g C}} = 1.00 \text{ moles C}$$



$$\text{H}_{8.00} \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 7.94 \text{ moles H}$$

$$\text{C}_{3.33} \text{H}_4 \times \frac{1 \text{ mole}}{12.01 \text{ g C}} = 3.33 \text{ moles C} \rightarrow \text{C}_3 \text{H}_4$$

Empirical Formula Pogil

Theodore Rudolph Running



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