

Experiment 7 Report Sheet

Empirical Formulas

Date _____ Lab Sec _____ Name _____ Desk No. _____
 Indicate whether the following data are for Part B, Part C, or Part D. Identify the original sample _____

	Trial 1	Trial 2
1. Mass of crucible and lid (g)	25.636	40.347
2. Mass of crucible, lid, and sample (g)	25.786	40.542
3. Mass of sample (g) <u>0.150</u>	0.150	0.195
4. Instructor's approval of flame and apparatus	—	—
5. Mass of crucible, lid, and product		
1st mass measurement (g)	25.858	40.654
2nd mass measurement (g)	25.856	40.652
3rd mass measurement (g)	25.859	40.658
6. Final mass of crucible, lid, and product (g)	25.860	40.655
7. Mass of product (g)	0.224g	0.113g
8. Part B. Combination Reaction of Magnesium and Oxygen		
a. Mass of O (g)	0.014g	
b. Mass ratio of Mg to O	2:1	
c. Moles of Mg (mol)	0.00617 mol	
d. Moles of O (mol)	0.00463 mol	
e. Mole ratio of Mg to O	1:1	
f. Consensus empirical formula of magnesium oxide	<u>1 MgO</u>	
g. Percent by mass (%): <u>67.0</u> % Mg; <u>33.0</u> % O		

9. Part C. Decomposition Reaction of a Pure Compound

a. Mass of product = mass CaO (g)	0.224	
b. Mass of CO ₂ (g)	0.1263	

e) note

Empirical Formula Of A Compound Lab Report

H Kauffman



Empirical Formula Of A Compound Lab Report:

Reviewing **Empirical Formula Of A Compound Lab Report**: Unlocking the Spellbinding Force of Linguistics

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Table of Contents Empirical Formula Of A Compound Lab Report

1. Understanding the eBook Empirical Formula Of A Compound Lab Report
 - The Rise of Digital Reading Empirical Formula Of A Compound Lab Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Empirical Formula Of A Compound Lab Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Empirical Formula Of A Compound Lab Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical Formula Of A Compound Lab Report
 - Personalized Recommendations
 - Empirical Formula Of A Compound Lab Report User Reviews and Ratings

- Empirical Formula Of A Compound Lab Report and Bestseller Lists
- 5. Accessing Empirical Formula Of A Compound Lab Report Free and Paid eBooks
 - Empirical Formula Of A Compound Lab Report Public Domain eBooks
 - Empirical Formula Of A Compound Lab Report eBook Subscription Services
 - Empirical Formula Of A Compound Lab Report Budget-Friendly Options
- 6. Navigating Empirical Formula Of A Compound Lab Report eBook Formats
 - ePub, PDF, MOBI, and More
 - Empirical Formula Of A Compound Lab Report Compatibility with Devices
 - Empirical Formula Of A Compound Lab Report Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Empirical Formula Of A Compound Lab Report
 - Highlighting and Note-Taking Empirical Formula Of A Compound Lab Report
 - Interactive Elements Empirical Formula Of A Compound Lab Report
- 8. Staying Engaged with Empirical Formula Of A Compound Lab Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Empirical Formula Of A Compound Lab Report
- 9. Balancing eBooks and Physical Books Empirical Formula Of A Compound Lab Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Empirical Formula Of A Compound Lab Report
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Empirical Formula Of A Compound Lab Report
 - Setting Reading Goals Empirical Formula Of A Compound Lab Report
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Empirical Formula Of A Compound Lab Report
 - Fact-Checking eBook Content of Empirical Formula Of A Compound Lab Report
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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