

Empirical Formula Practice 1.0 – Answer Key

Objective: - Calculate the empirical formula of a compound given experimental data

Directions: Answer the following empirical formula questions. Show all work! Include units!

1. Dibutyl succinate is an insect repellent used against household ants and roaches. Its composition is 62.58% C, 9.83% H and 27.79% O. Find its empirical formula. $[C_{16}H_{22}O_4]$

$$C: = 5.210658 \text{ mol} + 1.736875 = 3.0000 \quad = 3 \quad \times 2 = 6$$

$$H: = 9.534653 \text{ mol} + 1.736875 = 5.4895 \quad = 5.5 \quad \times 2 = 11$$

$$O: = 1.736875 \text{ mol} + 1.736875 = 1 \quad = 1 \quad \times 2 = 2$$



2. The compound adrenaline contains 56.7% C, 6.56% H, 8.28% N, and 28.37% O by mass. What is the empirical formula for adrenaline?

$$C: = 4.721066 \text{ mol} + 0.591006 = 7.9882 \quad = 8$$

$$H: = 6.495050 \text{ mol} + 0.591006 = 10.9698 \quad = 11$$

$$N: = 0.591006 \text{ mol} + 0.591006 = 1 \quad = 1$$

$$O: = 1.773125 \text{ mol} + 0.591006 = 3.0002 \quad = 3$$



3. A sample of urea contains 0.480 g of C, 0.161 g of H, 1.121 g of N, and 0.640 g of O. What is the empirical formula of urea?

$$C: = 0.039967 \text{ mol} + 0.039967 = 1 \quad = 1$$

$$H: = 0.159406 \text{ mol} + 0.039967 = 3.9664 \quad = 4$$

$$N: = 0.080014 \text{ mol} + 0.039967 = 2.0002 \quad = 2$$

$$O: = 0.040000 \text{ mol} + 0.039967 = 1.0008 \quad = 1$$



Empirical Formula Problems Answer Key

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Table of Contents Empirical Formula Problems Answer Key

1. Understanding the eBook Empirical Formula Problems Answer Key
 - The Rise of Digital Reading Empirical Formula Problems Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Empirical Formula Problems Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Empirical Formula Problems Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical Formula Problems Answer Key
 - Personalized Recommendations
 - Empirical Formula Problems Answer Key User Reviews and Ratings
 - Empirical Formula Problems Answer Key and Bestseller Lists

5. Accessing Empirical Formula Problems Answer Key Free and Paid eBooks
 - Empirical Formula Problems Answer Key Public Domain eBooks
 - Empirical Formula Problems Answer Key eBook Subscription Services
 - Empirical Formula Problems Answer Key Budget-Friendly Options
6. Navigating Empirical Formula Problems Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Empirical Formula Problems Answer Key Compatibility with Devices
 - Empirical Formula Problems Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Empirical Formula Problems Answer Key
 - Highlighting and Note-Taking Empirical Formula Problems Answer Key
 - Interactive Elements Empirical Formula Problems Answer Key
8. Staying Engaged with Empirical Formula Problems Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Empirical Formula Problems Answer Key
9. Balancing eBooks and Physical Books Empirical Formula Problems Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Empirical Formula Problems Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Empirical Formula Problems Answer Key
 - Setting Reading Goals Empirical Formula Problems Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Empirical Formula Problems Answer Key
 - Fact-Checking eBook Content of Empirical Formula Problems Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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