

Let the relative atomic mass of element Z be  $z$ .

Element	Z	Bromine, Br
Mass (g)	2.07	$3.67 - 2.07 = 1.60$
Number of moles of atoms	$\frac{2.07}{z}$	$\frac{1.60}{80} = 0.02$

Based on the empirical formula  $ZBr_2$ , the ratio of atoms of Z : Br is 1 : 2.

Therefore,  $\frac{2.07}{z} : 0.02 = 1 : 2$

$$\frac{2.07}{z} \div 0.02 = 1 \div 2$$

$$\frac{2.07}{0.02z} = \frac{1}{2}$$

$$z = \frac{2 \times 2.07}{0.02}$$

$$= 207$$

The relative atomic mass of element Z is 207.

# Empirical Formula For Copper Oxide

**VM Jensen**



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