

3) The empirical formula of a compound represents the lowest whole number mole ratio of component elements.

4) The molecular formula is the actual ratio of elements found within the naturally-occurring sample. This could also equal the empirical formula if the naturally-occurring formula is also in the lowest whole number mole ratio.

3) Molecular = $C_2H_6NO_2$, \leftarrow each subscript can be reduced by 2:

Empirical = CH_3NO

4) Compound pair B is the only set of compounds with the same empirical formulae.

5) To determine the empirical formula, convert each atom's provided mass into moles and find the lowest whole number mole ratio:

$$0.8007 \text{ g C} \times 1 \text{ mol C}/12.01 \text{ g C} = 0.06676 \text{ mol C}$$

$$0.9330 \text{ g H} \times 1 \text{ mol H}/1.008 \text{ g H} = 0.92663 \text{ mol H}$$

$$0.2006 \text{ g N} \times 1 \text{ mol N}/14.01 \text{ g N} = 0.02861 \text{ mol N}$$

divide by the smallest # of moles

$$2.0007 \text{ g O} \times 1 \text{ mol O}/16.00 \text{ g O} = 0.12503 \text{ mol O}$$

$$0.06676 \text{ mol C}/0.12503 = 0.53 \text{ mol C}$$

$$0.92663 \text{ mol H}/0.12503 = 7.41 \text{ mol H}$$

$$0.02861 \text{ mol N}/0.12503 = 0.23 \text{ mol N}$$

CH_3NO_2 = empirical formula for

carboxylic acid

$$0.12503 \text{ mol O}/0.12503 = 2 \text{ mol O}$$

6) Determine the empirical formula first, by assuming that 1.00g of the sample is available, following for 24.27% C \rightarrow 24.27g C, etc... Use the relative masses of each element to determine moles and find the lowest whole number mole ratio. Then compare the mass of the empirical formula to that of the mass of the molecular formula provided. This comparison will allow you to determine the molecular formula.

$$24.27\% \text{ C} \rightarrow 24.27 \text{ g C} \times 1 \text{ mol C}/12.01 \text{ g C} = 2.023 \text{ mol C}$$

$$4.07\% \text{ H} \rightarrow 4.07 \text{ g H} \times 1 \text{ mol H}/1.008 \text{ g H} = 4.037 \text{ mol H}$$

$$71.46\% \text{ Cl} \rightarrow 71.46 \text{ g Cl} \times 1 \text{ mol Cl}/35.45 \text{ g Cl} = 2.021 \text{ mol Cl}$$

$$2.023 \text{ mol C}/2.021 = 1 \text{ mol C}$$

$$4.037 \text{ mol H}/2.021 = 2 \text{ mol H}$$

$$2.023 \text{ mol Cl}/2.021 = 1 \text{ mol Cl}$$

CH_2Cl = empirical formula

Empirical mass of CH_2Cl \rightarrow C: 12.03 g +

H: 2(1.008g) +

Cl: 35.45 g

= 46.476 g

To compare empirical & molecular mass:

Molecular mass/empirical mass = n (n= empirical formula) \rightarrow molecular formula

$$96.95 \text{ g}/46.476 \text{ g} = 2$$

$2CH_2Cl$ = $C_2H_4Cl_2$

Empirical And Molecular Formulas Answer Key

Jicheng Xie

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