#### Multiple Choice Questions

- K<sub>1</sub> and K<sub>2</sub> are the velocity constants of forward and backward reactions.
  The equilibrium constant k of the reaction is
  - (A)  $K_j \times K_j$

(B)  $K_1 - K_2$ 

(C)  $\frac{K_I}{K_I}$ 

 $(D) \frac{K_I + K_I}{K_I - K_I}$ 

- Sed. (C).
- Which of the following reactions will be favoured at low pressure?
  - (A) N, +0, -2NO

- (B) H<sub>2</sub> + I<sub>2</sub> → 2HI
- (C)  $PCI_A \longrightarrow 2PCI_A + CI_A$
- (D)  $N_j + 3H_j = 5NH_j$

- Sed. (C).
- For the reaction, A+B--+C+D, if active mass of A is kept constant and active mass of B is tripled, the rate of reaction will become
  - (A) three times

(B) six times

(C) eight times

(D) nine times

- Sol. (D).
- 4. The rate of  $\frac{K_p}{K_s}$  for the reaction:

$$CO(g) + \frac{1}{2}O_{1}(g) = CO_{1}(g)$$
 is

(A), I,

(B) RT

 $(C) (RT)^{I/I}$ 

(D)  $(RT)^{-\epsilon}$ 

- Sol. (D).
- 5. Consider the reaction,

in a closed container at equilibrium. What would be the effect of addition of CaCO, on the equilibrium concentration of CO2?

(A) Increases

tBY Unpredictable

(C) Decreases:

(D) Remains unaffected

- Sed. (106.)
- For the reaction, PG<sub>2</sub>(g), PG<sub>2</sub>(g) + G<sub>2</sub>(g), the forward reaction at constant temperature is favoured by
  - (A) introducing an mert gas at constant volume
  - (B) introducing PCI(g) at constant volume
  - (C) introducing PCl+(g) at constant volume
  - (D) introducing Cl3(g) at constant volume
- Sol. (C).
- A liquid is in equilibrium with its vapour at its boiling point. On an average, the molecules in the two phases have equal
  - (A) intermolecules forces
  - (B) potential energy
  - (C) kimeric energy
  - (D) both kinetic energy and potential energy
- Soul. (C).
- 8. The unit of equilibrium constant for the reversible reaction

# **Multiple Choice Questions In Chemical Equilibrium**

**Arshad Iqbal** 

#### **Multiple Choice Questions In Chemical Equilibrium:**

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Unveiling the Energy of Verbal Artistry: An Mental Sojourn through Multiple Choice Questions In Chemical Equilibrium

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