

Chapter

7

THERMOCHEMISTRY

MULTIPLE CHOICE QUESTIONS

- Spontaneous reactions are: (a) Reversible (b) Irreversible (c) Non irreversible (d) None of these **(LHR 05)**
- The standard heat changes occur at: (a) 25°C and 2 atm (b) 298 K and 1 atm (c) 25°C and 1 mm hg (d) 273 K and 1 atm **(BWP 08, GRW 06)**
- $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$ the change in enthalpy for reaction is called: (a) Heat of reaction (b) heat of formation (c) Heat of neutralization (d) Heat of combustion **(LHR 06)**
- The net change in energy in a chemical reaction is same whether it takes place directly or indirectly. It is called: (a) Henry's law (b) Charlie's law (c) Hess's law (d) Graham's law **(RWP 11, GRW 07, BWP 10)**
- The enthalpy change when one mole of substance is completely burnt in excess of oxygen is called: (a) Enthalpy of atomization (b) Enthalpy of neutralization (c) Enthalpy of Combustion (d) Enthalpy of formation **(LHR 07)**
- Enthalpy change for the reaction: $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$ is called enthalpy of: (a) Formation (b) Combustion (c) Neutralization (d) Atomization **(LHR 08)**
- Standard enthalpies are measured at: (a) 273 K (b) 298K (c) 373 K (d) All of these **(FSD 11, LHR 09)**
- The exothermic process is: (a) Evaporation (b) Sublimation (c) Respiration (d) Boiling **(DGK 10, GRW, LHR 11)**
- A state function which describes together the internal energy and the product of pressure and volume is called: (a) Enthalpy (b) Internal energy (c) Work (d) Free energy **(GRW 11)**

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Neil D. Jespersen, Pamela Kerrigan



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