

1) Which of the following statements is incorrect?

- A) The degree of ionization of a weak electrolyte increases with dilution.
- B) Strong electrolytes are ionized completely at all concentrations.
- C) Addition of NH_4Cl to NH_4OH increases the ionization of the latter.
- D) Increase the temperature increases the ionization.

2) According to Lowry and Bronsted system the chloride ion (Cl^-) in aqueous solution is a

- A) Weak base
- B) Strong base
- C) Weak acid
- D) Strong acid

3) In the following reactions, ZnO is respectively acting as a/an

i) $\text{ZnO} + \text{Na}_2\text{O}$ gives Na_2ZnO_2

ii) $\text{ZnO} + \text{CO}_2$ gives ZnCO_3

- A) Acid and acid
- B) Acid and base
- C) Base and acid
- D) Base and base

4) Given K_{a1} , K_{a2} for $\text{H}_2\text{S} = 1.1 \times 10^{-7} \text{M}^2$. 1L of 0.1 M H_2S having $[\text{H}^+] = 0.1 \text{M}$ contains S^{2-} ions equal to.

- A) 6.62×10^3
- B) 6.62×10^4
- C) 6.62×10^5
- D) 6.62×10^6

5) Equal volume of three acid solutions of pH 3, 4 and 5 are mixed in a vessel. What will be

the pH of the resulting solution?

- A) 3.33
- B) 3.5
- C) 3.67
- D) 3.83

6) The pH of a solution is 4. The concentration of H^+ ions in the solution is

- A) 10^{-4}
- B) 10^{-5}
- C) 10^{-6}
- D) 10^{-7}

7) The degree of dissociation of H_2S is 0.1. The concentration of H^+ ions is

- A) 0.1
- B) 0.2

Multiple Choice Questions On Chemical Equilibrium

G Psacharopoulos



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