

1. Acetic acid in benzene solution forms dimer due to intermolecular H-bonding. For this case van't Hoff factor is:
- (A) $i = 1$ (B) $i > 1$
(C) $i < 1$ (D) inclusive
2. 90% of a first-order reaction completes in 90 minutes. 50% of the reaction will be over in approximately:
- (A) 50 minutes (B) 54 minutes
(C) 27 minutes (D) 62 minutes
3. Lanthanide contraction is caused due to
- (A) The appreciable shielding on outer electrons by 4f electrons from the nuclear charge
(B) The appreciable shielding on outer electrons by 5d electrons from the nuclear charge
(C) The same effective nuclear charge from Ce to Lu
(D) The imperfect shielding on outer electrons by 4f electrons from the nuclear charge

90% completion
 $\ln \frac{100}{10} = \frac{k \cdot 90}{2.303}$
 $\ln 10 = \frac{k \cdot 90}{2.303}$
 $k = \frac{2.303 \ln 10}{90}$
50% completion
 $\ln \frac{100}{50} = \frac{k \cdot t}{2.303}$
 $\ln 2 = \frac{k \cdot t}{2.303}$
 $t = \frac{2.303 \ln 2}{k}$
 $t = \frac{2.303 \ln 2}{\frac{2.303 \ln 10}{90}}$
 $t = \frac{90 \ln 2}{\ln 10}$
 $t = \frac{90 \times 0.693}{0.301}$
 $t = 207.9 \text{ minutes}$

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