

Mathematics B (Modular)													
Unit grade boundaries				Max Mark	a*	a	b	c	d	e	f	g	u
5MB1F	Number and algebra/Shape, space and measures 2 (Foundation)	Raw	60					49	41	33	25	17	0
		UMS	120					72	60	48	36	24	0
5MB1H	Number and algebra/Shape, space and measures 2 (Higher)	Raw	60	54	49	40	32	24	20			0	
		UMS	120	108	96	84	72	60	54			0	
5MB2F	Number, Algebra, Geometry 1 (Foundation) (Non-calculator)	Raw	60					54	45	36	28	20	0
		UMS	120					72	60	48	36	24	0
5MB2H	Number, Algebra, Geometry 1 (Higher) (Non-calculator)	Raw	60	54	48	38	28	18	13			0	
		UMS	120	108	96	84	72	60	54			0	
5MB3F	Number, Algebra, Geometry 2 (Foundation)	Raw	80					71	59	47	35	23	0
		UMS	160					96	80	64	48	32	0
5MB3H	Number, Algebra, Geometry 2 (Higher)	Raw	80	72	64	53	42	31	25			0	
		UMS	160	144	128	112	96	80	72			0	
Cash-in grade boundaries				Max Mark	A*	A	B	C	D	E	F	G	U
2MB01	Mathematics B	UMS	400	360	320	280	240	200	160	120	80	0	

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l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} \rightarrow \text{MnO}_4^-$
4 a a q is

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the following reduction oxidation reaction in basic solution what is the coefficient in front of OH^- $\text{Br}_2 + \text{I}^- \rightarrow \text{Br}^- + \text{I}_2$ previous
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web google classroom for the following redox equation $\text{MnO}_4^- + \text{SO}_2 + \text{H}_2\text{O} \rightarrow \text{Mn}^{2+} + \text{HSO}_4^-$ what will be the
stoichiometric coefficient of HSO_4^- if the above equation is balanced in an acidic medium stuck use a hint report a problem
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redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} \rightarrow \text{AuCN}_2^- + 2\text{e}^-$ $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ make electrons equal
convert to basic solution $4\text{CN}^- + 2\text{Au} \rightarrow 2\text{AuCN}_2^- + 2\text{e}^-$ multiplied by a factor of 2 $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance
being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all
reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h₂o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$ 2 equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a factor of 6

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding H^+ H_2O and e^- in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6$ $2\text{H}_2\text{NO}_2$ $\text{C}_6\text{H}_6\text{O}_6$ $2\text{H}_2\text{O}$ 2NO choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 b $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 c $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 d $\text{C}_6\text{H}_8\text{O}_6$ and NO_2

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