



Name _____ Class _____ Date _____

1

How many atoms of **oxygen (O)** are in the **chemical formula** below?

- A** 4
- B** 5
- C** 7
- D** 12

**3**

Carbon (C) and hydrogen (H) create **covalent bonds** to share **electrons** between them. What is the total number of electrons being shared between **carbon** and **hydrogen** in the molecule of propane shown below?

- A** 8
- B** 11
- C** 16
- D** 32

**2**

The diagram shown below is a **structural formula** for the gas **propane**. What would be the correct **chemical formula** for propane?

- A** C_3H_6
- B** CH_6
- C** $8\text{H}_3\text{C}$
- D** CH

**4**

The **reason** why the bonds between carbon (C) and hydrogen (H) are **covalent** is _____

- A** C and H do not have electrons to give away
- B** C and H are both nonmetals
- C** C and H are both metalloids
- D** C and H can bond only by sharing electrons

**5**

The diagram below represents the **structural formula** for several **water molecules**. How many **molecules** are represented?

- A** 3
- B** 5
- C** 10
- D** 15

**6**

Explain the **arrangement** of the molecules in the diagram of water below.

- A** it is a lucky arrangement
- B** water molecules are sticky
- C** water molecules have polarity
- D** water molecules are magnetic

**7**

Water molecules have some **polarity** in their structure. Using the diagram below, notice how the ends of H atoms (**positive**) and O atoms (**negative**) meet. This property is the **reason** that _____

- A** water forms droplets easily
- B** water boils easily
- C** ice melts easily
- D** water has color

**8**

Since H atoms are slightly **positive** and the O atoms are slightly **negative**, molecules of water have **polarity**. This makes water a **good** _____

- A** solute
- B** refrigerant
- C** solvent
- D** magnet

**9**

The chemical formula for the **oxygen** that we breathe is _____

- A** O
- B** O_2
- C** O_3
- D** 2O

**10**

How many **molecules of oxygen** are represented in the chemical formula below?

- A** 2
- B** 4
- C** 6
- D** 18



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RS Peters



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