

PROBLEM 2-37

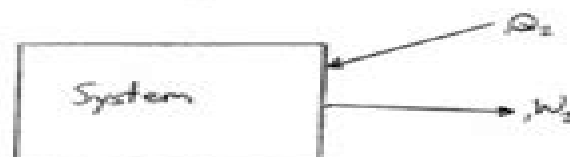
GIVEN: A system of fixed mass at constant pressure and temperature

FIND: Equivalent system form of Eqn. 2.35 (First Law) which is used to define the heat of reaction.

ASSUMPTIONS: expansion/contraction of the system boundary is a reversible process; only work done by system is boundary expansion/contraction

APPROACH: Write the first law for the system. Substitute the appropriate expression for constant pressure work and solve for the specific heat transfer.

Sketch of system:



first law for a system:

$$Q_2 - W_2 = m(u_2 - u_1)$$

for a system at constant pressure the reversible work can be expressed as

$$W_2 = \int P dV = P(V_2 - V_1) = mP(v_2 - v_1) \quad v \equiv \text{specific volume}$$

substituting into the first law

$$Q_2 - mP(v_2 - v_1) = m(u_2 - u_1)$$

$$Q_2 = m[(u_2 + Pv_2) - (u_1 + Pv_1)] = m(h_2 - h_1)$$

$$q_2 = Q_2/m = h_2 - h_1$$

if state 1 is the reactants and state 2 is the products then

$$q = h_p - h_r \quad \leftarrow \text{same form as Eqn 2.35}$$

COMMENTS: Note that the first law for a constant pressure system has the same form as for a control volume. This is only true for constant pressure. Also, the internal energies u_1 and u_2 are mixture internal energies. Even assuming ideal gas behavior, $u_1 \neq u_2$ at constant temperature if the composition changes. (see Joule's experiment with gases and internal energy)

Introduction To Combustion Turns Solution Manual

BM King



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Table of Contents Introduction To Combustion Turns Solution Manual

1. Understanding the eBook Introduction To Combustion Turns Solution Manual
 - The Rise of Digital Reading Introduction To Combustion Turns Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Introduction To Combustion Turns Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Introduction To Combustion Turns Solution Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Introduction To Combustion Turns Solution Manual
 - Personalized Recommendations
 - Introduction To Combustion Turns Solution Manual User Reviews and Ratings
 - Introduction To Combustion Turns Solution Manual and Bestseller Lists
5. Accessing Introduction To Combustion Turns Solution Manual Free and Paid eBooks

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- Introduction To Combustion Turns Solution Manual Budget-Friendly Options
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 - Introduction To Combustion Turns Solution Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Introduction To Combustion Turns Solution Manual
 - Highlighting and Note-Taking Introduction To Combustion Turns Solution Manual
 - Interactive Elements Introduction To Combustion Turns Solution Manual
- 8. Staying Engaged with Introduction To Combustion Turns Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Introduction To Combustion Turns Solution Manual
- 9. Balancing eBooks and Physical Books Introduction To Combustion Turns Solution Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Introduction To Combustion Turns Solution Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Introduction To Combustion Turns Solution Manual
 - Setting Reading Goals Introduction To Combustion Turns Solution Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Introduction To Combustion Turns Solution Manual
 - Fact-Checking eBook Content of Introduction To Combustion Turns Solution Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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