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Joseph Devlin



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buffers typically consist of an acid base pair with the acid and base differing by the presence or absence of a proton a conjugate acid base pair for instance one of the buffers that maintain the pH of human blood involves carbonic acid H_2CO_3 and its conjugate base the bicarbonate ion HCO_3^-

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an acid and a base react not to produce a salt and a solvent but to form a new acid and a new base the concept of neutralization is thus absent brønsted lowry acid base behavior is formally independent of any solvent making it more all encompassing than the arrhenius model

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arrhenius's definition of acids and bases the earliest definition of acids and bases is arrhenius's definition which states that an acid is a substance that forms hydrogen ions H^+ when dissolved in water and a base is a substance that forms hydroxide ions OH^- when dissolved in water

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the reaction of an acid and a base is called a neutralization reaction although acids and bases have their own unique chemistries the acid and base cancel each other's chemistry to produce a rather innocuous substance water in fact the general reaction between an acid and a base is acid + base → water + salt

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