

### Specific Heat Problems – ANSWERS

The last page of this document has Miscellaneous Information that you will need for these problems.

1. How many calories are needed to warm 55.8 g of water at 21.3°C to 73.4°C?

$$\text{Heat} = m \times S. H. \times (T_{\text{final}} - T_{\text{initial}})$$

$$\text{Heat} = 55.8 \text{ g} \times 1.00 \text{ cal/g}^\circ\text{C} \times (73.4^\circ\text{C} - 21.3^\circ\text{C})$$

$$\text{Heat} = 55.8 \text{ g} \times 1.00 \text{ cal/g}^\circ\text{C} \times 52.1^\circ\text{C}$$

$$\text{Heat} = 2907.18 \text{ cal} \quad \text{convert to scientific notation}$$

$$= 2.9\textbf{0}718 \times 10^3 \text{ cal}$$

$$= 2.91 \times 10^3 \text{ cal (3 sig. figs.)}$$

2. How many calories are needed to warm 55.8 g of alcohol from 21.3°C to 73.4°C?

$$\text{Heat} = m \times S. H. \times (T_{\text{final}} - T_{\text{initial}})$$

$$\text{Heat} = 55.8 \text{ g} \times 0.58 \text{ cal/g}^\circ\text{C} \times (73.4^\circ\text{C} - 21.3^\circ\text{C})$$

$$\text{Heat} = 55.8 \text{ g} \times 0.58 \text{ cal/g}^\circ\text{C} \times 52.1^\circ\text{C}$$

$$\text{Heat} = 1686.16 \text{ cal} \quad \text{convert to scientific notation}$$

$$\text{Heat} = 1.6\textbf{8}616 \times 10^3 \text{ cal}$$

$$\text{Heat} = 1.69 \times 10^3 \text{ cal}$$

# Heat Practice Problems Answers

**Dave Prowse**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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American Steam and Hot-water Heating Practice Engineering Record, Building Record and Sanitary Engineer, 1895  
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