

Specific Heat and Heat Capacity

Name _____ KEY _____ Period _____

1. Calculate the amount of heat needed to increase the temperature of 250 g of water from 20 °C to 46 °C.

$$\begin{aligned}q &= m \times C \times \Delta T \\q &= 250 \text{ g} \times 4.18 \text{ J/g}^\circ\text{C} \times 26^\circ\text{C} \\q &= 27,170 \text{ J or } 27 \text{ kJ}\end{aligned}$$

2. Calculate the specific heat capacity of copper given that 204.75 J of energy raises the temperature of 15 g of copper from 25 °C to 60. °C.

$$\begin{aligned}q &= m \times C \times \Delta T \\C &= q/m \times \Delta T \\C &= 204.75 \text{ J} / (15 \text{ g} \times 35^\circ\text{C}) \\C &= 0.39 \text{ J/g}^\circ\text{C}\end{aligned}$$

3. 216 J of energy is required to raise the temperature of aluminum from 15 °C to 35 °C. Calculate the mass of aluminum. (Specific heat capacity of aluminum is 0.90 J/g °C).

$$\begin{aligned}q &= m \times C \times \Delta T \\m &= q/C \times \Delta T \\m &= 216 \text{ J} / (0.90 \text{ J/g}^\circ\text{C} \times 20^\circ\text{C}) \\m &= 12 \text{ g}\end{aligned}$$

4. The initial temperature of 150 g of ethanol was 22 °C. What will be the final temperature of the ethanol if 3240 J was needed to raise the temperature of the ethanol? (Specific heat capacity of ethanol is 2.44 J/g °C).

$$\begin{aligned}q &= m \times C \times \Delta T \\\Delta T &= q/m \times C \\\Delta T &= 3240 \text{ J} / (150 \text{ g} \times 2.44 \text{ J/g}^\circ\text{C}) \\\Delta T &= 8.85^\circ\text{C} \\T_{\text{final}} &= 22^\circ\text{C} + 8.85^\circ\text{C} = 31^\circ\text{C}\end{aligned}$$

5. When 435 J of heat is added to 3.4 g of olive oil at 21 °C, the temperature increases to 85 °C. What is the specific heat of the olive oil?

$$\begin{aligned}q &= m \times C \times \Delta T \\C &= q/m \times \Delta T \\C &= 435 \text{ J} / (3.4 \text{ g} \times 64^\circ\text{C}) \\C &= 2.0 \text{ J/g}^\circ\text{C}\end{aligned}$$

6. How many kilojoules of heat are absorbed when 1.00 L of water is heated from 18 °C to 85 °C? (Hint: You first need to determine the mass of the water, then calculate q in the requested unit.)

Heat Calculation Answer Key

Jin-Ying Zhang



Heat Calculation Answer Key:

The Enigmatic Realm of **Heat Calculation Answer Key**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing in short supply of extraordinary. Within the captivating pages of **Heat Calculation Answer Key** a literary masterpiece penned by a renowned author, readers attempt a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting effect on the hearts and minds of those that partake in its reading experience.

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