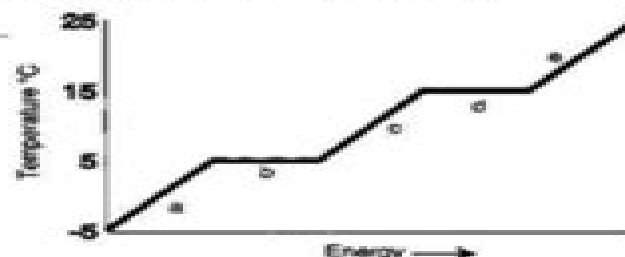


HEATING CURVES QUESTIONS

PART A – HEATING CURVES. Use the heating curve to the right to answer the following questions:

1. What is the melting point of the substance? **B/ 5 C**
2. What is the boiling point of the substance? **D/ 15 C**
3. Which letter represents heating of the solid? **A**
4. Which letter represents heating of the gas? **E**
5. Which letter represents melting of the solid? **B**
6. Which letter represents boiling of the liquid? **D**
7. At points b and d why doesn't the temperature change even though heat energy is added?

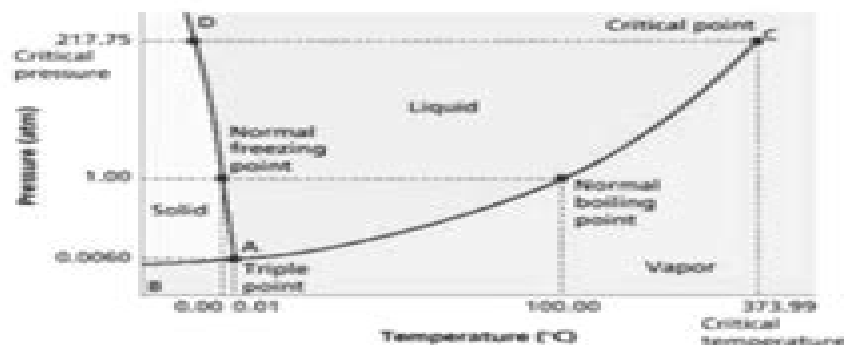


The Substance must absorb enough heat/ kinetic energy to Escape the intermolecular actions/ bonds (i.e. hydrogen bonds) that Keep them their current phase/ state

Refer to the phase diagram below when answering the questions on this worksheet

PART C – PHASE DIAGRAMS. Use the phase diagram for water below to answer the following questions.

8. What is the state of water at 2 atm and 50°? **LIQUID**
9. What phase change will occur if the temperature is lowered from 80°C to -5°C at 1 atm? **FREEZING**
10. You have ice at -10°C and 1 atm. What could you do in order cause the ice to sublime? **Decrease the pressure to less than 0.060 atm**



Predict the physical state of each of these substances at the indicated temperature. Use the melting point and boiling point data from the table below.

- a. phenol at 99°C
- b. ammonia at -25°C
- c. methanol in an ice-water bath
- d. methanol in a boiling-water bath
- e. ammonia at -100°C
- f. phenol at 25°C

- a. **Liquid**
- b. **Gas**
- c. **Liquid**
- d. **Gas**
- e. **Solid**
- f. **Solid**

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Jon L. Dossett, Robert E. Luetje



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