

## → CHEMICAL EQUILIBRIUM

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**Q.1:** Justify that chemical equilibrium is dynamic in nature?

**Ans:** In reversible chemical reactions the molecules of reactions collide and convert into products. At the same time the molecules of the products are converted into reactants. When two opposing forces maintain the equal, rates then equilibrium is there and that is dynamic equilibrium in nature.

**Q.2:** When does four types of chemical equilibrium constants for a reaction become equal?

**Ans:** When we have such a reversible reaction, in which the total number of moles of reactants and products are equal then  $K_p = K_c = K_n$ . In other words  $\Delta n = 0$ , where  $\Delta n$  is the difference of number of moles of reactants and products.

**Q.3:** The change of temperature disturbs the equilibrium position and the equilibrium constant of reaction? Justify?

**Ans:** All the reversible reactions are disturbed by changing their equilibrium position and equilibrium constant by disturbing the temperature. Actually change of temperature changes the energy contents of reactants and products.

**Q.4:** Why is solubility of glucose in water is increased by increasing the temperature?

**Ans:** The solubility of glucose in water is an endothermic process. Increase of temperature pushed the system to that side where heat is absorbed, and that is the side where solubility increases.

**Q.5:** What is the nature of solution having  $\text{pH} = 7.5$ ?

**Ans:** The solution with  $\text{pH} = 7$  is neutral and having values more than 7 have greater concentration of  $\text{OH}^-$  and less concentration of  $\text{H}^+$ , so the solution is basic.

**Q.6:** What is buffer capacity?

**Ans:** Buffer capacity is a ability of a buffer to resist the change of  $\text{pH}$ , when a few drops of an acid and base is added from outside. The best buffer is obtained. (i) when  $\text{pH}$  of the buffer is equal to  $\text{pK}_a$  of the acid and the concentration of the salt and acid are equal. (ii) The  $\text{OH}^-$  of the buffer is equal to  $\text{pK}_b$  or the base.

# Multiple Questions On Chemical Equilibrium

**Madan R.L.**



## Multiple Questions On Chemical Equilibrium:

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## **Table of Contents Multiple Questions On Chemical Equilibrium**

1. Understanding the eBook Multiple Questions On Chemical Equilibrium
  - The Rise of Digital Reading Multiple Questions On Chemical Equilibrium
  - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Questions On Chemical Equilibrium
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Multiple Questions On Chemical Equilibrium
  - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Questions On Chemical Equilibrium
  - Personalized Recommendations
  - Multiple Questions On Chemical Equilibrium User Reviews and Ratings
  - Multiple Questions On Chemical Equilibrium and Bestseller Lists
5. Accessing Multiple Questions On Chemical Equilibrium Free and Paid eBooks
  - Multiple Questions On Chemical Equilibrium Public Domain eBooks
  - Multiple Questions On Chemical Equilibrium eBook Subscription Services
  - Multiple Questions On Chemical Equilibrium Budget-Friendly Options
6. Navigating Multiple Questions On Chemical Equilibrium eBook Formats
  - ePub, PDF, MOBI, and More
  - Multiple Questions On Chemical Equilibrium Compatibility with Devices
  - Multiple Questions On Chemical Equilibrium Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Multiple Questions On Chemical Equilibrium
  - Highlighting and Note-Taking Multiple Questions On Chemical Equilibrium
  - Interactive Elements Multiple Questions On Chemical Equilibrium
8. Staying Engaged with Multiple Questions On Chemical Equilibrium

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Multiple Questions On Chemical Equilibrium
- 9. Balancing eBooks and Physical Books Multiple Questions On Chemical Equilibrium
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Multiple Questions On Chemical Equilibrium
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Multiple Questions On Chemical Equilibrium
  - Setting Reading Goals Multiple Questions On Chemical Equilibrium
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Multiple Questions On Chemical Equilibrium
  - Fact-Checking eBook Content of Multiple Questions On Chemical Equilibrium
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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