

Grade 9

Paper 1

Time : 1 ½ hours

Marks :70

Examiner :

Date :

Instructions:

Read the following instructions carefully before answering the questions.

1. This EXAMINATION consists of SIX questions and covers the following topics:

SECTION A

Short Questions	Question 1	10 marks
	Question 2	6 marks
	Question 3	5 marks
	<u>Total section A:</u>	<u>21 marks</u>

SECTION B:

Accounting Equation	Question 4	10 marks
Journals	Question 5	31 marks
General Ledger	Question 6	8 marks
	<u>Total section B:</u>	<u>49 marks</u>

TOTAL 70 marks

2. Answer ALL the questions.
 3. Read the instructions for each question carefully and take note of what is required.
 4. Number the answers carefully according to the numbering system used in this question paper. No marks will be awarded for answers that are numbered incorrectly.
 5. Use the mark allocation below as a guide when answering each question.
 6. Write neatly and legibly.
 7. Good luck!
-

Maths Paper Scope For Final Exams

**Trung Tran,Cuong Huu Nguyen,Loc Thi
My Nguyen**

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calculate the number of moles from each of the following a 6.95×10^{23} atoms of Cu b 7.12×10^{21} molecules of CO₂ c 3.33×10^{22} ions of Pb⁴⁺ d 7.00×10^{23} formula units of LiNO₃

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the quantity of heat for a process is represented by the letter q the sign of q for an endothermic process is positive because the system is gaining heat a chemical reaction or physical

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one milligram is equal to 0 001 grams 790 drag each number to the correct location sort the numbers based on the number of significant figures they have one significant figure 1 000

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the quantities commonly encountered in chemical problems include the number of moles of a substance the number of atoms molecules or formula units of a substance and the mass in

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it is always a good idea to estimate the answer before you do the actual calculation in this case the mass given 35 00 g is less than the molar mass so the answer should be less than 1

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