

Organic Chemistry

PETROLEUM = FOSSIL FUEL

Petroleum - dead organisms that fell to ocean floor, buried under thick sediment
High pressure slowly converted them to petroleum

Fossil fuel are petroleum, coal and natural gas
- the remaining of plants and animals

Natural gas - methane
Often found with petroleum
High temperature and pressure cause the compounds to break down to gas

Coal - remains of lush vegetation buried under thick sediment
Pressure and heat slowly converted into coal

What is petroleum?

- A smelly mixture of hundreds of different compounds
- Organic compound - mostly hydrocarbons
- A non-renewable resource
- Refine petroleum, carried out by fractional distillation to separate compounds into groups with molecules of similar size

Carbon atoms bond to each other
hydrogen bond to carbon atoms



Fractions

Boiling point	Flow	Volatile	Easily burn?	Size of molecules
up to 100°C	very runny	volatile	very easy	small
100°C - 150°C	runny	less volatile	easy	↓ large
150°C - 200°C	not very runny	even less volatile	not easy	
200°C - 300°C	viscous (thick and sticky)	least volatile	only with a wick	

TRENDS

- The larger the molecules in a hydrocarbon
 - the higher its boiling point
 - the less volatile
 - the less easy to flow / more viscous
 - the less easy to burn

Refinery gas	1-4 carbon atoms	bottled gases, cooking, heating	25°C
Gasoline (petrol)	5-6 carbon atoms	fuel for cars	boiling point and viscosity ↓ 400°C
Naphtha	6-10 carbon atoms	feedstock for chemicals and plastics	
Paraffin (kerosene)	10-15 carbon atoms	fuel for aircraft, oil stoves, lamps	
Diesel oil (gas oil)	15-20 carbon atoms	fuel for diesel engines	
Fuel oil	20-30 carbon atoms	fuel for power stations, ships, home heating system	
Lubricating fraction	30-50 carbon atoms	oil for car engines, machinery, motors, polishes	
Bitumen	> 50 carbon atoms	for road surfaces and roofs	

Igcse Organic Chemistry Study Guide

S Ashworth



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