Chemistry Revision

Moles, Empirical formula, Ar and Mr

A. S. M.

- Relative atomic mass:
 - the weighted average mass of the naturally occurring isotopes of an element as compared to 1/12 of the mass of carbon-12
 - weighted average means it factors in the abundance of each isotope.
 - (% x A) + (% x A) / 100
- Relative formula mass:
 - the sum of all the atomic masses.
 - use relative formula mass to work out percentage composition.

Question: find percentage by mass of copper in copperfit oxide, CuO

$$M_2 = 64 + 16$$

Percentage of copper = 64/80 x 100

$$-.80\%$$

Miches:

- The number of carbon atoms in exactly 12g of carbon-12.
- Epomulaci
 - number of moles = mass (g) / M.
 - number of moles = number of particles / 6.02 x 10.23
- Reacting mass calculations:
 - mole (convert mass into moles);
 - mole ration (from balanced equation):
 - un-mole (convert back to grams)
 - when un-moling in gas calculations, always multiply by 24 instead of M.

Question: find the mass aluminium oxide when 135g of aluminium is burned with oxygen. 4At + 3O₀ → 2A₀O₀

- mole = mass/M = 135/27 = 5 mol Al
- mole ratio = 4:2 = 2:1 = 2.5 mol Al₂O₂
- un-mole = 2.5 x 102 = 255g Al₂O₅

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