

1 This question is about reversible reactions.

Sulfur trioxide is produced in a reversible reaction.

The equation for the reaction is:



1.1 What is the correct balanced equation for the reaction?

Tick **one** box.

- ☐ $2\text{SO}_2 + \text{O}_2 \rightleftharpoons \text{SO}_3$
- ☐ $\text{SO}_2 + 2\text{O} \rightleftharpoons \text{SO}_3$
- ☐ $2\text{SO}_2 + \text{O}_2 \rightleftharpoons 2\text{SO}_3$
- ☐ $2\text{SO}_2 + 2\text{O}_2 \rightleftharpoons 2\text{SO}_3$

[1 mark]

1.2 The reaction of sulfur dioxide and oxygen to form sulfur trioxide is exothermic.

Which of the following statements is true?

Tick **one** box.

- ☐ The activation energy for the reaction of sulfur dioxide with oxygen is negative.
- ☐ Sulfur trioxide is colder than sulfur dioxide and oxygen.
- ☐ Heating the reaction will cause more sulfur trioxide to form.
- ☐ The reaction of sulfur trioxide to form sulfur dioxide and oxygen is endothermic.

[1 mark]

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1. Understanding the eBook Igcse Chemistry Past Papers Aqa
 - The Rise of Digital Reading Igcse Chemistry Past Papers Aqa
 - Advantages of eBooks Over Traditional Books
2. Identifying Igcse Chemistry Past Papers Aqa
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Igcse Chemistry Past Papers Aqa
 - User-Friendly Interface
4. Exploring eBook Recommendations from Igcse Chemistry Past Papers Aqa
 - Personalized Recommendations

- Igcse Chemistry Past Papers Aqa User Reviews and Ratings
- Igcse Chemistry Past Papers Aqa and Bestseller Lists
- 5. Accessing Igcse Chemistry Past Papers Aqa Free and Paid eBooks
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 - Igcse Chemistry Past Papers Aqa eBook Subscription Services
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 - Igcse Chemistry Past Papers Aqa Enhanced eBook Features
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 - Interactive Elements Igcse Chemistry Past Papers Aqa
- 8. Staying Engaged with Igcse Chemistry Past Papers Aqa
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Igcse Chemistry Past Papers Aqa
- 9. Balancing eBooks and Physical Books Igcse Chemistry Past Papers Aqa
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Igcse Chemistry Past Papers Aqa
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Igcse Chemistry Past Papers Aqa
 - Setting Reading Goals Igcse Chemistry Past Papers Aqa
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Igcse Chemistry Past Papers Aqa
 - Fact-Checking eBook Content of Igcse Chemistry Past Papers Aqa

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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