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(Campbell, Bush, Brunell, 2005)

# How To Reference A In A Report

**Andy Field, Graham Hole**



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 partly because so many compounds can act as acids or bases

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 a 25 0 ml sample of 0 100 m hydrochloric acid with 0 100 m sodium hydroxide the values of the ph measured after successive  
 additions of small amounts of naoh are listed in the first column of this table and are graphed in

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 until the solution is exactly neutralised i e there is only salt and water it is possible to calculate

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 titration of an acid with a base or vice versa the indicator is a substance that can exist in two forms an

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web the above equation works only for neutralizations in which there is a 1 1 ratio between the acid and the base the  
 example below demonstrates the technique to solve a titration problem for a titration of sulfuric acid with sodium hydroxide

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